

Chemistry
Higher level
Paper 1

Thursday 11 May 2017 (afternoon)

1 hour

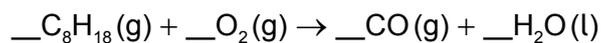
Instructions to candidates

- Do not open this examination paper until instructed to do so.
- Answer all the questions.
- For each question, choose the answer you consider to be the best and indicate your choice on the answer sheet provided.
- The periodic table is provided for reference on page 2 of this examination paper.
- The maximum mark for this examination paper is **[40 marks]**.

The Periodic Table

	1	2	3	4	5	6	7	8	9	10	11	12	13	14	15	16	17	18																																																									
1	1 H 1.01	<table border="1"> <tr> <td colspan="19">Atomic number</td> </tr> <tr> <td colspan="19">Element</td> </tr> <tr> <td colspan="19">Relative atomic mass</td> </tr> </table>																	Atomic number																			Element																			Relative atomic mass																		
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2	3 Li 6.94	4 Be 9.01																																																																									
3	11 Na 22.99	12 Mg 24.31																																																																									
4	19 K 39.10	20 Ca 40.08	21 Sc 44.96	22 Ti 47.87	23 V 50.94	24 Cr 52.00	25 Mn 54.94	26 Fe 55.85	27 Co 58.93	28 Ni 58.69	29 Cu 63.55	30 Zn 65.38	31 Ga 69.72	32 Ge 72.63	33 As 74.92	34 Se 78.96	35 Br 79.90	36 Kr 83.90																																																									
5	37 Rb 85.47	38 Sr 87.62	39 Y 88.91	40 Zr 91.22	41 Nb 92.91	42 Mo 95.96	43 Tc (98)	44 Ru 101.07	45 Rh 102.91	46 Pd 106.42	47 Ag 107.87	48 Cd 112.41	49 In 114.82	50 Sn 118.71	51 Sb 121.76	52 Te 127.60	53 I 126.90	54 Xe 131.29																																																									
6	55 Cs 132.91	56 Ba 137.33	57 † La 138.91	72 Hf 178.49	73 Ta 180.95	74 W 183.84	75 Re 186.21	76 Os 190.23	77 Ir 192.22	78 Pt 195.08	79 Au 196.97	80 Hg 200.59	81 Tl 204.38	82 Pb 207.2	83 Bi 208.98	84 Po (209)	85 At (210)	86 Rn (222)																																																									
7	87 Fr (223)	88 Ra (226)	89 † Ac (227)	104 Rf (267)	105 Db (268)	106 Sg (269)	107 Bh (270)	108 Hs (269)	109 Mt (278)	110 Ds (281)	111 Rg (281)	112 Cn (285)	113 Uut (286)	114 Uug (289)	115 Uup (288)	116 Uuh (293)	117 Uus (294)	118 Uuo (294)																																																									
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	58 Ce 140.12	59 Pr 140.91	60 Nd 144.24	61 Pm (145)	62 Sm 150.36	63 Eu 151.96	64 Gd 157.25	65 Tb 158.93	66 Dy 162.50	67 Ho 164.93	68 Er 167.26	69 Tm 168.93	70 Yb 173.05	71 Lu 174.97																																																													
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	90 Th 232.04	91 Pa 231.04	92 U 238.03	93 Np (237)	94 Pu (244)	95 Am (243)	96 Cm (247)	97 Bk (247)	98 Cf (251)	99 Es (252)	100 Fm (257)	101 Md (258)	102 No (259)	103 Lr (262)																																																													

1. What is the sum of the coefficients when the equation is balanced with whole numbers?



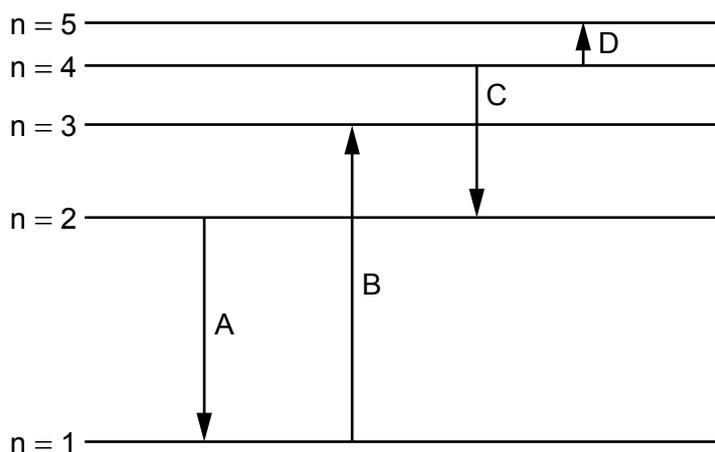
- A. 26.5
B. 30
C. 53
D. 61
2. What is the maximum volume, in dm^3 , of $\text{CO}_2(\text{g})$ produced when 1.00 g of $\text{CaCO}_3(\text{s})$ reacts with 20.0cm^3 of 2.00mol dm^{-3} $\text{HCl}(\text{aq})$?



Molar volume of gas = $22.7\text{ dm}^3\text{ mol}^{-1}$; $M_r(\text{CaCO}_3) = 100.00$

- A. $\frac{1}{2} \times \frac{20.0 \times 2.00}{1000} \times 22.7$
B. $\frac{20.0 \times 2.00}{1000} \times 22.7$
C. $\frac{1.00}{100.00} \times 22.7$
D. $\frac{1.00}{100.00} \times 2 \times 22.7$
3. Which factors affect the molar volume of an ideal gas?
- I. Pressure
II. Temperature
III. Empirical formula
- A. I and II only
B. I and III only
C. II and III only
D. I, II and III

4. Which electron transition emits radiation of the longest wavelength?



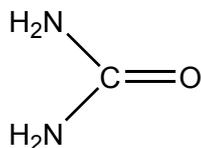
5. X, Y and Z represent the successive elements, Ne, Na and Mg, but not necessarily in that order.

	First ionization energy / kJ mol^{-1}
X	2081
Y	496
Z	738

What is the order of increasing atomic number?

- A. $X < Y < Z$
- B. $X < Z < Y$
- C. $Y < Z < X$
- D. $Y < X < Z$
6. Which property increases down Group 1, the alkali metals?
- A. Atomic radius
- B. Electronegativity
- C. First ionization energy
- D. Melting point

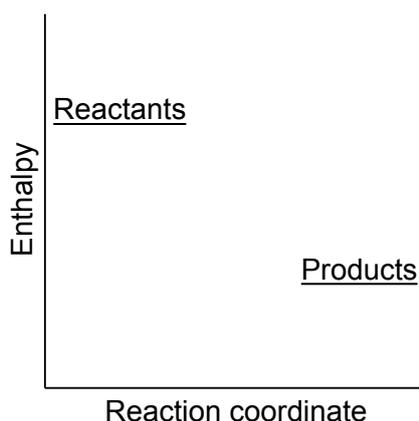
7. Which element is a lanthanide?
- A. Hf
 - B. Tb
 - C. U
 - D. Y
8. Ammonia is a stronger ligand than water. Which is correct when concentrated aqueous ammonia solution is added to dilute aqueous copper(II) sulfate solution?
- A. The d-orbitals in the copper ion split.
 - B. There is a smaller splitting of the d-orbitals.
 - C. Ammonia replaces water as a ligand.
 - D. The colour of the solution fades.
9. How many bonding electrons are there in the urea molecule?



- A. 8
 - B. 16
 - C. 20
 - D. 24
10. Which does **not** show resonance?
- A. PO_4^{3-}
 - B. C_6H_6
 - C. C_6H_{12}
 - D. O_3

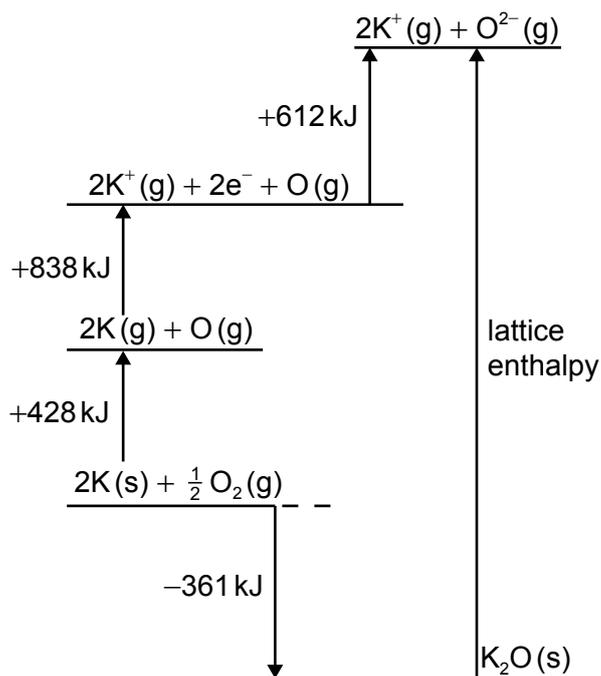
11. Which metal has the strongest metallic bond?
- A. Li
 - B. Na
 - C. K
 - D. Rb
12. Which is the first step in the CFC-catalysed destruction of ozone in UV light?
- A. $\text{CCl}_2\text{F}_2 \rightarrow \text{CClF}_2^+ + \text{Cl}^-$
 - B. $\text{CCl}_2\text{F}_2 \rightarrow \cdot\text{CClF}_2 + \text{Cl}\cdot$
 - C. $\text{CCl}_2\text{F}_2 \rightarrow \text{CCl}_2\text{F}^+ + \text{F}^-$
 - D. $\text{CCl}_2\text{F}_2 \rightarrow \cdot\text{CCl}_2\text{F} + \text{F}\cdot$
13. Which statement is correct?
- A. Sigma bonds are formed only by the combination of s atomic orbitals.
 - B. Pi bonds can be formed in the absence of sigma bonds.
 - C. Pi bonds are formed parallel to the axis between atoms.
 - D. Pi bonds are formed only by the combination of hybrid orbitals.

14. What can be deduced from this reaction profile?



- A. The reactants are less stable than the products and the reaction is exothermic.
 - B. The reactants are less stable than the products and the reaction is endothermic.
 - C. The reactants are more stable than the products and the reaction is exothermic.
 - D. The reactants are more stable than the products and the reaction is endothermic.
15. What can be deduced from the facts that ozone absorbs UV radiation in the region of 340 nm and molecular oxygen in the region of 242 nm?
- A. The bond between atoms in molecular oxygen is a double bond.
 - B. The bonds in ozone are delocalized.
 - C. The bonds between atoms in ozone are stronger than those in molecular oxygen.
 - D. The bonds between atoms in molecular oxygen need more energy to break.

16. The Born-Haber cycle for potassium oxide is shown below:



Which expression represents the lattice enthalpy in kJ mol^{-1} ?

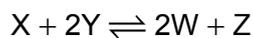
- A. $-361 + 428 + 838 + 612$
 - B. $-(-361) + 428 + 838 + 612$
 - C. $-361 + 428 + 838 - 612$
 - D. $-(-361) + 428 + 838 - 612$
17. Which ion's hydration energy is the most exothermic?
- A. Li^+
 - B. Na^+
 - C. Br^-
 - D. I^-

Questions 18 and 19 refer to the following reaction.



18. Which change does **not** increase the initial rate of reaction when $\text{CaCO}_3(\text{s})$ is added to excess $\text{HCl}(\text{aq})$?
- A. Decrease in the size of the $\text{CaCO}_3(\text{s})$ particles
 - B. Increase in the temperature of the reaction mixture
 - C. Increase in the concentration of $\text{HCl}(\text{aq})$, keeping the same volume
 - D. Increase in the volume of $\text{HCl}(\text{aq})$, keeping the same concentration
19. Which methods can be used to monitor the progress of this reaction?
- I. Change in colour of this reaction mixture
 - II. Change in mass of this reaction mixture
 - III. Change in volume of gas evolved
- A. I and II only
 - B. I and III only
 - C. II and III only
 - D. I, II and III
20. Which is true of an Arrhenius plot of $\ln k$ (y -axis) against $\frac{1}{T}$?
- A. The graph goes through the origin.
 - B. The activation energy can be determined from the gradient.
 - C. The intercept on the x -axis is the activation energy.
 - D. The intercept on the y -axis is the frequency factor, A .

21. Which is correct about reaction mechanisms?
- A. A species that is zero order does not take part in the reaction.
 - B. A catalyst does not take part in the reaction.
 - C. Reactants in a fast step before the slow step are included in the rate expression.
 - D. Reactants in a fast step after the slow step are included in the rate expression.
22. Which variable affects the equilibrium constant, K_c ?
- A. Atmospheric pressure
 - B. Catalyst
 - C. Concentration of reactants
 - D. Temperature
23. Components X and Y are mixed together and allowed to reach equilibrium. The concentrations of X, Y, W and Z in the equilibrium mixture are 4, 1, 4 and 2 mol dm^{-3} respectively.



What is the value of the equilibrium constant, K_c ?

- A. $\frac{1}{8}$
- B. $\frac{1}{2}$
- C. 2
- D. 8

24. Which of the following does **not** react with dilute HCl(aq)?

Extract from activity series

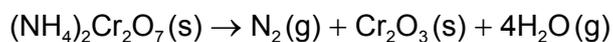


- A. Na_2CO_3
 - B. Cu
 - C. Zn
 - D. CuO
25. Which of the following is correct?
- A. A weak acid is a proton donor and its aqueous solution shows good conductivity.
 - B. A weak acid is a proton donor and its aqueous solution shows poor conductivity.
 - C. A weak acid is a proton acceptor and its aqueous solution shows good conductivity.
 - D. A weak acid is a proton acceptor and its aqueous solution shows poor conductivity.
26. Which type of bond is formed when a Lewis acid reacts with a Lewis base?
- A. Covalent
 - B. Dipole-dipole
 - C. Double
 - D. Hydrogen

27. What is the order of increasing acidity of the following acids?

Acid	K_a	Acid	pK_a
chloroethanoic	1.3×10^{-3}	hydrogen fluoride	3.3
ethanoic	1.7×10^{-5}	hydrogen cyanide	9.3

- A. chloroethanoic < ethanoic < hydrogen fluoride < hydrogen cyanide
- B. ethanoic < chloroethanoic < hydrogen fluoride < hydrogen cyanide
- C. chloroethanoic < ethanoic < hydrogen cyanide < hydrogen fluoride
- D. hydrogen cyanide < ethanoic < hydrogen fluoride < chloroethanoic
28. Which element is reduced in the following decomposition?



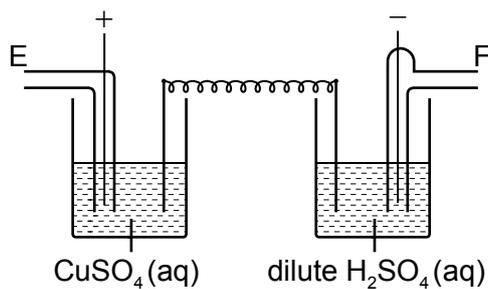
- A. N
- B. H
- C. Cr
- D. O
29. Which of the following is **not** a redox reaction?

- A. $\text{CH}_4(\text{g}) + \text{Cl}_2(\text{g}) \rightarrow \text{CH}_3\text{Cl}(\text{g}) + \text{HCl}(\text{g})$
- B. $\text{C}(\text{s}) + \text{O}_2(\text{g}) \rightarrow \text{CO}_2(\text{g})$
- C. $2\text{CO}(\text{g}) \rightarrow \text{CO}_2(\text{g}) + \text{C}(\text{s})$
- D. $\text{CH}_3\text{COOH}(\text{aq}) + \text{NaOH}(\text{aq}) \rightarrow \text{CH}_3\text{COONa}(\text{aq}) + \text{H}_2\text{O}(\text{l})$

30. What is the standard half-cell potential of copper if the “zero potential reference electrode” is changed from the standard hydrogen electrode to a standard zinc electrode?

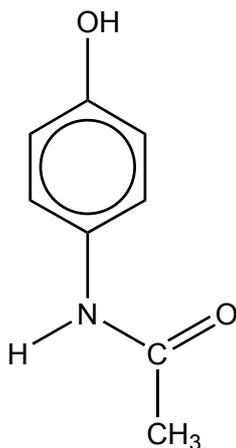
	E^\ominus / V with respect to the standard hydrogen electrode
$\text{Zn}^{2+}(\text{aq}) + 2\text{e}^- \rightleftharpoons \text{Zn}(\text{s})$	-0.76
$\text{Cu}^{2+}(\text{aq}) + 2\text{e}^- \rightleftharpoons \text{Cu}(\text{s})$	+0.34

- A. -1.1
 B. -0.34
 C. +0.34
 D. +1.1
31. What are the relative volumes of gas given off at E and F during electrolysis of the two cells in series? Assume all electrodes are inert.



- A. 1:1
 B. 1:2
 C. 2:1
 D. 5:2

32. Which functional group is present in paracetamol?



- A. Carboxyl
- B. Amino
- C. Nitrile
- D. Hydroxyl

33. Which describes the reaction between a halogen and ethane?

	Mechanism	Bond fission in halogen
A.	free radical	homolytic
B.	free radical	heterolytic
C.	addition	homolytic
D.	addition	heterolytic

34. Which compound contains a secondary carbon atom?

- A. $\text{CH}_3\text{CH}(\text{Cl})\text{CH}(\text{CH}_3)_2$
- B. $(\text{CH}_3)_2\text{CHCH}_2\text{Cl}$
- C. $(\text{CH}_3)_3\text{CCl}$
- D. $\text{CH}_3\text{CH}_2\text{Cl}$

35. Which pair of isomers always shows optical activity?

- A. Cis-trans
- B. Enantiomers
- C. Conformational
- D. E/Z

36. Which compounds can be reduced?

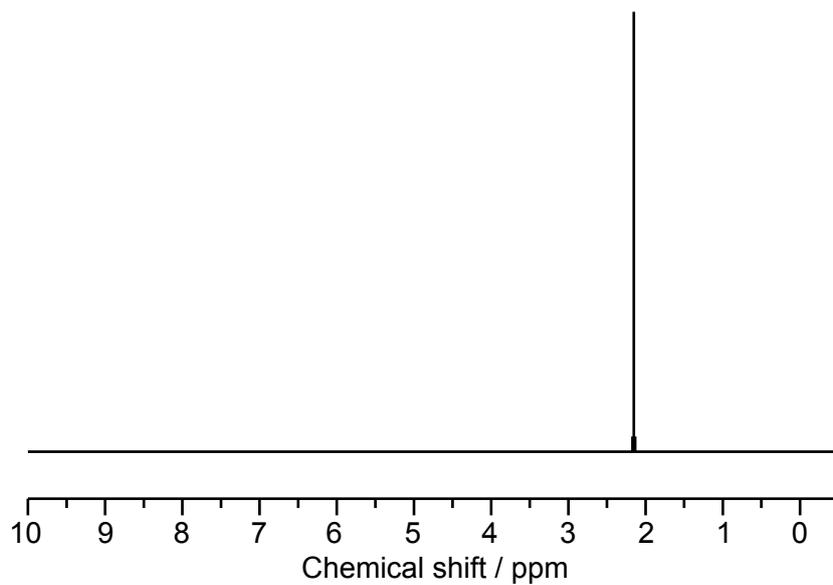
- I. C_2H_4
- II. CH_3COOH
- III. CH_3CHO

- A. I and II only
- B. I and III only
- C. II and III only
- D. I, II and III

37. In which order should the reagents be used to convert benzene into phenylamine (aniline)?

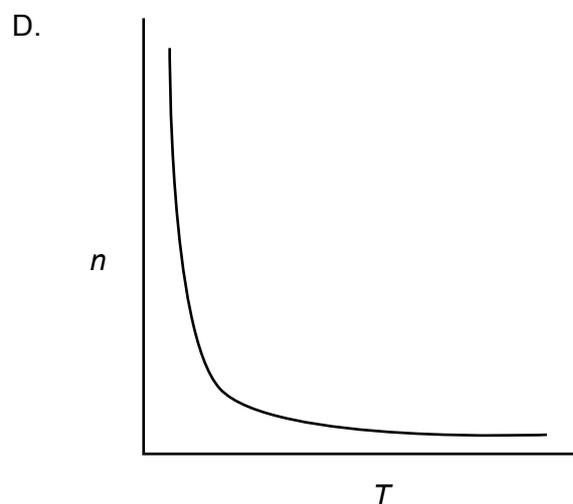
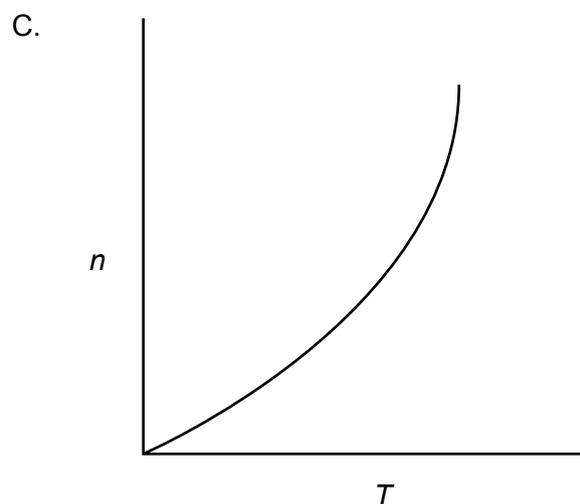
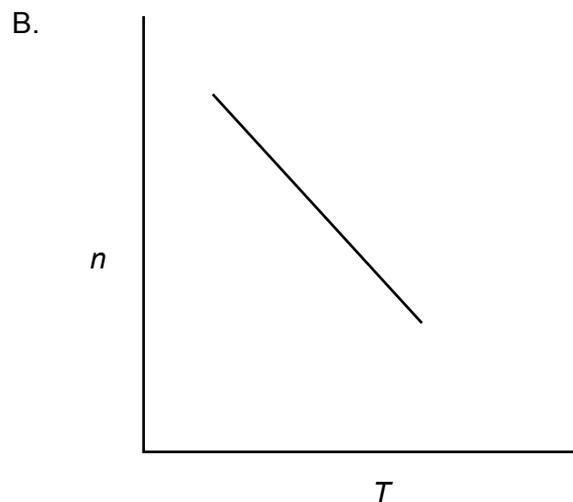
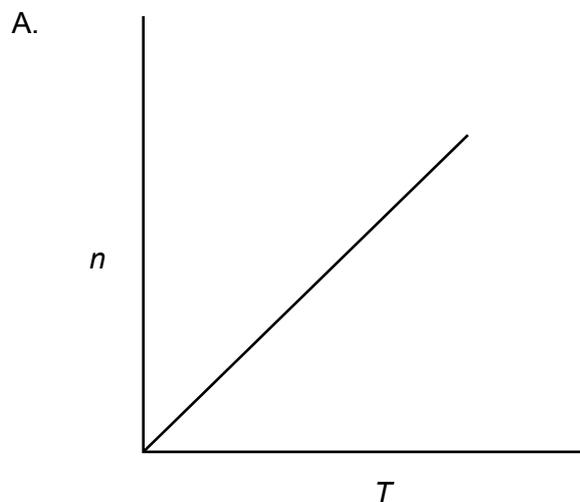
	1st reagent	2nd reagent	3rd reagent
A.	NaOH	Sn / conc. HCl	conc. HNO_3 / conc. H_2SO_4
B.	Sn / conc. HCl	NaOH	conc. HNO_3 / conc. H_2SO_4
C.	conc. HNO_3 / conc. H_2SO_4	Sn / conc. HCl	NaOH
D.	NaOH	conc. HNO_3 / conc. H_2SO_4	Sn / conc. HCl

38. What can be deduced from the following ^1H NMR spectrum?



- A. There is only one hydrogen atom in the molecule.
- B. There is only one hydrogen environment in the molecule.
- C. The molecule is a hydrocarbon.
- D. There is only one isotope in the element.

39. What is the graphical relationship between n and T in the ideal gas equation, $pV = nRT$, all other variables remaining constant?



40. Which technique can be used to identify bond length and bond angle?

- A. ^1H NMR spectroscopy
 - B. IR spectroscopy
 - C. Mass spectroscopy
 - D. X-ray crystallography
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